

## Writing and Balancing Chemical Equations

For each of the following write a **balanced molecular equation** and an **ionic equation** to show the reaction that occurs, and then an **observation** for the reaction.

1. Some copper sulphate solution is added to some sodium hydroxide solution. /3
  2. Some lead (IV) nitrate solution is added to some sodium chloride solution. /3
  3. Sulphuric acid is added to some solid calcium carbonate. /3
  4. Hydrochloric acid is added to some solid calcium hydroxide. /3
  5. Copper sulphate solution is added to sodium carbonate solution. /3
  6. Magnesium sulphate solution is added to some barium nitrate solution. /3
  7. Potassium chloride solution is added to some sodium sulphate solution. /3
  8. Silver nitrate and barium chloride solutions are mixed. /3
  9. Nitric acid and sodium bicarbonate solutions are added to each other. /3
  10. Zinc chloride and sodium hydroxide solutions are mixed. /3
  11. Zinc sulfide is heated in air. /3
  12. Potassium metal is added to water. /3
  13. Potassium phosphate solution is added to magnesium chloride solution. /3
  14. Tin (II) bromide solution is added to a solution of ammonium sulphide. /3
  15. Lead (II) nitrate solution is mixed with potassium iodide solution. /3
- TOTAL /45