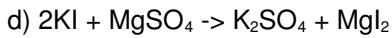
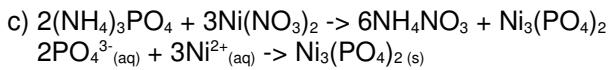
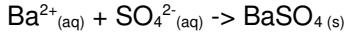
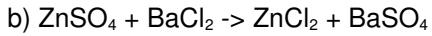
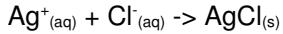
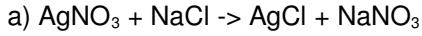


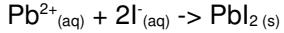
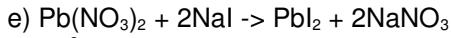
1.

- a) $\text{CaF}_2 \text{ (s)}$
- b) $\text{NaI} \text{ (aq)}$
- c) $\text{MgS} \text{ (s)}$
- d) $\text{ZnCO}_3 \text{ (s)}$
- e) $\text{PbCl}_4 \text{ (s)}$
- f) $(\text{NH}_4)_2\text{CO}_3 \text{ (aq)}$
- g) $\text{AgNO}_3 \text{ (aq)}$
- h) $\text{K}_3\text{PO}_4 \text{ (aq)}$
- i) $\text{Al(OH)}_3 \text{ (s)}$
- j) $\text{FeSO}_4 \text{ (aq)}$

2.



No reaction



3.

Mix a sample with an iodide – if there's a precipitate, it's lead

Otherwise mix a sample with a hydroxide – if there's a precipitate, it's magnesium

Otherwise it's sodium

4.

- 1 K
- 2 Ag
- 3 Ba
- 4 Cu
- 5 Sr
- 6 Fe^{2+}
- 7 Pb
- 8 Na