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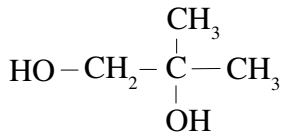	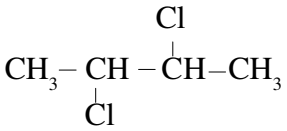
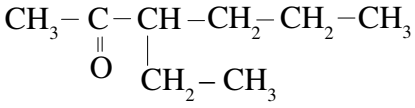	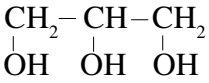
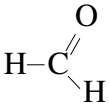	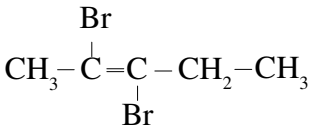
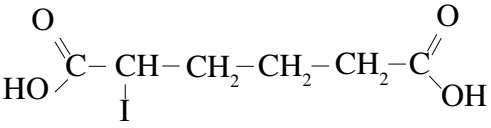	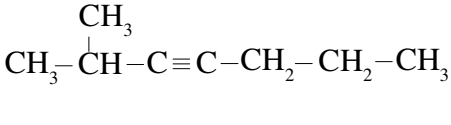
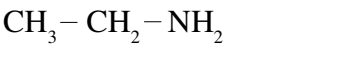	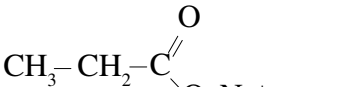

2.

a) 5 methyl heptanal (2 marks)	d) N, N dimethyl ethanamine (3 marks)
b) N ethyl 2-butanamine (3 marks)	e) sodium butanoate (2 marks)
c) propyl propanoate (2 marks)	f) N,5 dimethyl N ethyl 2-hexanamine (4 marks)

3. By boiling. Hexane has a lower boiling point than decane, because it has a shorter carbon chain and therefore is not as strongly held by dispersion forces.

4.
   a) The compounds have the same molecular formula (the same number of each element)
   b) Boiling point. Isomeric acids have higher boiling points than isomeric esters because they have the OH bond that ester's don't. This allows them to exhibit hydrogen bonding, which is a strong secondary interaction.

5. Ethanoic acid has a carboxyl group which means it has a C=O bond as well as the OH bond which propan-1-ol has. This means there are more sites for hydrogen bonding, which means forces between molecules is higher, hence higher m. pt. and b. pt.