Chapter 5 Vocab Review NAME\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

|  | **Definition** | **Diagrams, analogies, questions, or notes** |
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| **Atomic number** | Defines the element. The number of protons in an atom. |  |
| **Periods** | The horizontal rows in the periodic table. Represents electron shells in the atom. |  |
| **Groups** | The vertical columns in the periodic table. Represents number of valence electrons. |  |
| **Valence shell** | The outermost electron shell. An atom is most stable when this is full. |  |
| **Electrons** | Small, negatively charged particles which orbit around the nucleus of an atom. |  |
| **Protons** | Positively charged particles in the nucleus of an atom. |  |
| **Neutrons** | Neutral (no charge) particles in the nucleus of an atom. |  |
| **Valence electrons** | Outermost electrons. Atoms transfer or share these during chemical reactions. |  |
| **Atomic mass** | The number of protons plus neutrons in an atom. |  |
| **Alkali metals** | The elements in Group 1. These react strongly with water or acids. |  |
| **Electron shell** | Layers around the atom that electrons are found in. Also known as energy levels or orbitals. |  |
| **Electron configuration** | The number of electrons in each electron shell of an atom. |  |