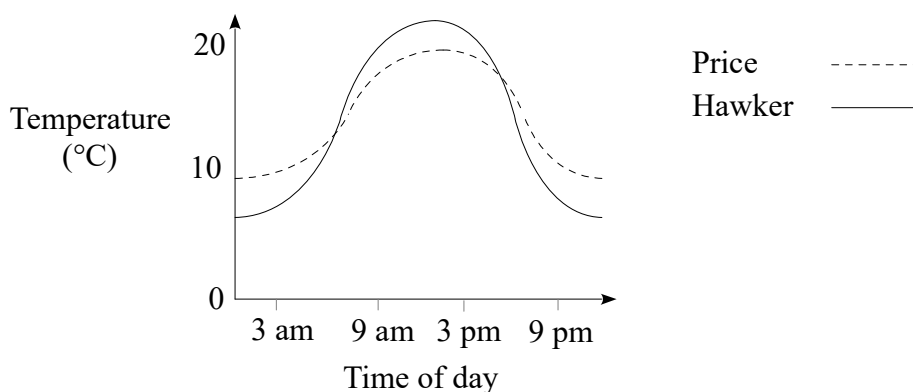


Heat

1.
 - (a) State one similarity and one difference between heat and temperature. /2
 - (b) Suggest and explain one way temperature can be measured without a thermometer. /3
2.
 - (a) Consider a metal object and a plastic object at the same temperature. Explain which one feels colder to touch. /2
 - (b) "During the winter, the insulation in our homes helps keep the cold out." Explain whether you agree or disagree with this statement. /2
 - (c) Use an example to explain how convection transfers heat energy. /2
 - (d) Write two questions for which the answer could include "Radiant heat". /2
3. Explain one example of how thermal expansion is used or worked around. /3
4.
 - (a) Use an analogy to explain thermal equilibrium. Thermal equilibrium is like... /2
 - (b) "A hot drink at 90°C will reach room temperature faster than a hot drink at 70°C." Explain whether you agree or disagree with this statement. /2
5. The graph below compares the average daily temperatures for the month of September in two South Australian towns. Price is a coastal town, Hawker is inland.



Explain why the difference between maximum and minimum temperatures is larger for Hawker than for Price. /2

6. The specific heat capacity of water is $4.18 \times 10^3 \text{ J/kg/K}$.
 - (a) Calculate the energy needed to heat 500 g of water by 50 K. /2
 - (b) When 50g of aluminium at 100°C is placed in water and cools to 30°C, it loses 3075 J of heat energy. Calculate the specific heat capacity of aluminium. /3
 - (c) Consider 10.0g of water at 100°C. If 1.0g of water evaporates, calculate the temperature of the remaining 9.0 grams. The latent heat of vaporisation for water is $2.272 \times 10^6 \text{ J/kg}$. /3
7.
 - (a) A car engine operates at an efficiency of about 25%. Explain what this means. /2
 - (b) An 11 W compact fluorescent lamp bulb produces about 0.99 W of light. Calculate the efficiency of the lamp. /2